

U.S.N.

# B.M.S. College of Engineering, Bengaluru-560019

Autonomous Institute Affiliated to VTU

## January / February 2025 Semester End Main Examinations

**Programme: B.E.**

**Branch: Chemical Engineering**

**Course Code: 19CH7DCBCE**

**Course: Biochemical Engineering**

**Semester: VII**

**Duration: 3 hrs.**

**Max Marks: 100**

**Instructions:** 1. Answer any FIVE full questions, choosing one full question from each unit.  
2. Missing data, if any, may be suitably assumed.

Important Note: Completing your answers, compulsorily draw diagonal cross lines on the remaining blank pages. Revealing of identification, appeal to evaluator will be treated as malpractice.			<b>UNIT - I</b>	<b>CO</b>	<b>PO</b>	<b>Marks</b>
	1	a)	With neat diagram, explain the structure of prokaryotic and eukaryotic cells.	CO1	PO2	06
		b)	Explain the structure of bacteria. Discuss the process of gram reaction to bacteria with suitable examples.	CO1	PO2	08
		c)	What are the monomers involved in the Deoxyribonucleic acid polymer? Explain the double helix of a Deoxyribonucleic acid.	CO2	PO2	06
			<b>OR</b>			
	2	a)	Brief about the significance of bioprocess engineering and give its application in various domain.	CO1	PO2	08
		b)	Explain in detail the classification of microorganisms.	CO1	PO2	12
			<b>UNIT - II</b>			
	3	a)	Explain how enzyme catalyzed reaction is different from chemical reaction? Describe the dependency of enzyme functionality with respect to pH, temperature and shear force.	CO4	PO2	10
		b)	When glucose is converted to fructose by glucose isomerase, the slow product formation step is also reversible as: $S + E \leftrightarrow ES$ ; where the rate constants are $k_1$ and $k_2$ respectively $ES \leftrightarrow P + E$ ; here the rate constants are $k_3$ and $k_4$ respectively Derive the rate equation considering Michaelis Menten and Briggs Haldane methods.	CO5	PO4	10
			<b>OR</b>			

4	a)	Explain the enzyme substrate complex formation for the enzyme catalyzed reaction. How does it lead to the basic expression for rate of enzyme catalyzed reaction?	CO5	PO4	8																											
	b)	<p>The data obtained for two different enzyme concentrations for enzyme catalyzed reaction:</p> <table><tr><th>Rate when <math>E_0 = 0.015</math> g/L</th><th>Substrate concentration in g/L</th><th>Rate When <math>E_0 = 8.75 \times 10^{-3}</math> g/L</th></tr><tr><td>1.14</td><td>20.0</td><td>0.67</td></tr><tr><td>0.87</td><td>10.0</td><td>0.51</td></tr><tr><td>0.7</td><td>6.7</td><td>0.41</td></tr><tr><td>0.59</td><td>5.0</td><td>0.34</td></tr><tr><td>0.5</td><td>4.0</td><td>0.29</td></tr><tr><td>0.44</td><td>3.3</td><td>-</td></tr><tr><td>0.39</td><td>2.9</td><td>-</td></tr><tr><td>0.35</td><td>2.5</td><td>-</td></tr></table> <p>Using Hanes-Woolf method, find i) <math>k_M</math> ii) Maximum rate for two different initial enzyme concentrations.</p>	Rate when $E_0 = 0.015$ g/L	Substrate concentration in g/L	Rate When $E_0 = 8.75 \times 10^{-3}$ g/L	1.14	20.0	0.67	0.87	10.0	0.51	0.7	6.7	0.41	0.59	5.0	0.34	0.5	4.0	0.29	0.44	3.3	-	0.39	2.9	-	0.35	2.5	-	CO5	PO4	12
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		UNIT - III																														
5	a)	From equilibrium approach, show that in the competitive enzyme inhibition the rate constant gets changed due to presence of inhibitors.	CO5	PO4	08																											
	b)	<p>In the following reaction of enzyme and substrate in presence of inhibitor, determine the rate expression for product formation.</p> $E + S \xrightleftharpoons{K_M} ES \xrightarrow{K_2} E + P$ $E + I \xrightleftharpoons{K_I} EI + S \leftrightarrow ESI$ $ES + I \xrightleftharpoons{K_{MI}} ESI$	CO5	PO4	12																											
		OR																														
6	a)	<p>The initial reaction rate of hydrolysis of a substrate by enzyme in absence and presence of inhibitor of <math>1.5 \times 10^{-7}</math> mol/L and the following data is obtained:</p> <table><tr><th>S mol/L</th><th>Initial reaction rate (no inhibitor)</th><th>Initial reaction rate (Inhibitor)</th></tr><tr><td>0.0032</td><td>0.111</td><td>0.059</td></tr><tr><td>0.0049</td><td>0.148</td><td>0.071</td></tr><tr><td>0.0062</td><td>0.143</td><td>0.091</td></tr><tr><td>0.0080</td><td>0.166</td><td>0.111</td></tr><tr><td>0.0095</td><td>0.2</td><td>0.125</td></tr></table>	S mol/L	Initial reaction rate (no inhibitor)	Initial reaction rate (Inhibitor)	0.0032	0.111	0.059	0.0049	0.148	0.071	0.0062	0.143	0.091	0.0080	0.166	0.111	0.0095	0.2	0.125	CO5	PO4	12									
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		i. Is it a case of competitive or non-competitive inhibition? ii. Evaluate Michaelis Menten kinetic parameters in the presence of inhibitor by Langmuir Plot.			
	b)	Explain the chemical method for enzyme immobilization.	CO4	PO2	08
		<b>UNIT - IV</b>			
7	a)	Explain the details of Monod growth kinetics. Also explain wash out in the said growth kinetics.	CO5	PO4	10
	b)	With a neat illustration, demonstrate the phases of cell growth.	CO3	PO4	04
	c)	Derive that the cell number doubling time is inversely proportional to the specific growth rate.	CO3	PO4	06
		<b>OR</b>			
8	a)	Derive the expression to obtain for specific growth rate at maximum cell output.	CO3	PO4	14
	b)	What is washout condition in growth kinetics? Explain.	CO3	PO4	6
		<b>UNIT - V</b>			
9	a)	Explain with diagram, the operation of stirred tank aseptic fermenter.	CO6	PO7	10
	b)	Explain the following downstream operations: i) Freeze drying ii) Affinity chromatography	CO6	PO7	10
		<b>OR</b>			
10	a)	Elaborate on aseptic aerobic fermenter operation.	CO6	PO7	10
	b)	Brief on medium formulation in bioreactor.	CO6	PO7	05
	c)	Explain the cell disruption using French press.	CO6	PO7	05

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